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**Activity One: Complete the table.**

Element	Atomic Number	Mass Number	Protons	Electrons	Neutrons
Helium					
Boron					
Sulfur					
Sodium					
Neon					
Hydrogen					
Calcium					

**Activity Two: Periodic table darts**

Calculate the total score for the throw (three darts).

Use a periodic table to find the element that matches the total – the total = atomic number.

DART 1	DART 2	DART 3	TOTAL	ELEMENT	SYMBOL
D3	17	5			
T5	T5	20			
D15	T7	18			
T1	T4	4			
BULL	18	1			
D20	T8	2			
14	16	19			

**Question II:**

1. Classify from the list given below as element, molecule and compound.

Pd | Si | MgCl<sub>2</sub> | H<sub>2</sub>O | Fe | SO<sub>2</sub> | CaH<sub>2</sub> | Sulfur | Br | O | NaOH | Ni | C | Oxygen | Li

- a) In the list given above, do you find substances that do not fit in the above categories? Name those substances. Give reasons for your answer.
- b) Write the name of all the compounds mentioned in your list.
- c) Mention the elements present in them; the list.

2. Complete the table given below –

	Element	Symbol	Z	A	p	n	s	Molar
a)	Phosphorus				15	15	15	
b)	Silicon				14	15		
c)								
d)								
e)	Argon							

3. Complete the table given below –

	Cation/Anion	Proton:nucleon	Chemical Formula	Name of the compound
a)	Al <sup>3+</sup>			Aluminum phosphate
b)		SO <sub>4</sub> <sup>2-</sup>		Ammonium sulphate
c)			CaCO <sub>3</sub>	
d)				Iron(II) chloride
e)				Magnesium carbonate

**Question III: Write the chemical formulae of the following compounds.**

1.	Ammonium chloride	N <sub>3</sub>	Chlorine hydride	
2.	Chlorine iodide	I <sub>2</sub>	Hydrogen iodide	
3.	Hydrogen bromide	HBr	Hydrogen chloride	
4.	Hydrogen cyanide	H <sub>2</sub> CN	Hydrogen fluoride	
5.	Hydrogen phosphate	HPO <sub>4</sub> <sup>2-</sup>	Hydrogen iodide	
6.	Lithium nitrate	LiNO <sub>3</sub>	Hydrogen carbide	
7.	Hydrogen sulfide	H <sub>2</sub> S	Hydrogen iodide	
8.	Hydrogen fluoride	HF	Hydrogen phosphate	
9.	Ammonium iodide	NH <sub>4</sub> I	Cyanogen carbide	
10.	Chlorine nitride	Cl <sub>2</sub> N	Hydrogen iodide	
11.	Hydrogen iodide	H <sub>2</sub> I	Hydrogen iodide	
12.	Sodium phosphate	Na <sub>3</sub> PO <sub>4</sub>	Hydrogen iodide	
13.	Copper carbide	Cu <sub>2</sub> C	Hydrogen iodide	

**Atomic Structure Worksheet 1**

Use the name examples below to help you name the atoms and ions. See if you can figure out the name differences between cation names and anion names or use your textbook for further help.

Use the information given on table below to calculate atomic mass. If not enough information is given on the table below then go to the periodic table to find the atomic mass. Round all atomic masses to the nearest whole number.

Nuclear Symbol	Atom/ ion name	Atom/ion Symbol	Number of Protons	Number of Neutrons	Number of Electrons	Atomic Mass	Atomic Number
	Nitrogen	N					
		C <sup>4</sup>					
		Os					
	Barium ion	Ba <sup>2+</sup>					
		Pb <sup>4+</sup>					
		Hf					

**Subatomic Particles**

*Students filled in the answers to these questions for homework assigned on WARM-UP QUIZ Thursday, 09/06, and checked their answers on Friday, 09/07.*

- What are the three subatomic particles?
- Where are the particles located in the atom?
- What are the charges of the particles?
- What does amu stand for?
- What is the mass (in amu) of each particle?
- Which of the subatomic particles is the lightest?
- What is the charge of the nucleus?
- Where is virtually all of the mass of the atom located?
- What effect do protons have on each other?
- What effect do electrons have on each other?
- What keeps the electrons in the atom?
- What is the symbol for each particle?
- What is the charge of an atom?
- What does a charge of an atom tell us about the number of protons and electrons?
- How is the nucleus of a hydrogen atom different from the nuclei of other elements?
- The most common form (isotope) of hydrogen does not have neutrons

**Writing Activity #2** Due Date: Monday, 09/09/2012

✓ Copy and include the symbol, location, charge, and relative mass (in amu) of the three subatomic particles. Label the table Document A.

✓ In a well-developed writing, compare and contrast the three subatomic particles. Use the table that you created as a supporting document and refer to it in your writing.

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